

Date.....

Name.....

Group

Lab report from practical lesson in biochemistry

Topic: Inorganic reactions

Wherever possible, provide the **net ionic equation** for the performed reaction. Record all your observations (formation of precipitate and its color, solution coloration etc.).

1.1 Reactions of silver compounds

a)

After addition of ammonia in excess:

b)

c)

After addition of ammonia in excess:

d)

1.2 Reactions of lead compounds

a)

b)

1.3 Reactions of copper compounds

a)

b)

After addition of ammonia in excess:

c)

d)

1.4 Reactions of mercury compounds

After addition of iodide in excess:

After addition of ammonia:

1.5 Reactions of iron compounds

a)

b)

c)

1.6 Reactions of calcium compounds

After acidification:

1.7 Reactions of ammonium cations

2. Flame reactions

Principle of flame reactions:

3.1 Reactions of sulfates

3.2 Reactions of carbonates and hydrogen carbonates

Heating of hydrogen carbonate:

Reaction of gas with limewater:

Reaction of the solid remnant after heating with acid:

3.3 Reactions of phosphates